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Here is an updated version of the \$domain website which many of our East European book trade customers have been using for some time now, more or less regularly. We have just introduced certain upgrades and changes which should be interesting for you. Please remember that our website does not replace publisher websites, there would be no point in

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duplicating the information. Our idea is to present you with tools that might be useful in your work with individual, institutional and corporate customers. Many of the features have been introduced at specific requests from some of you. Others are still at preparatory stage and will be implemented soon.

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question asks us to find the pH of the solution, so we will need to convert pOH to pH. To do so, we simply subtract the pOH from 14. The pH of the solution is 12.3.

Because sodium hydroxide is a strong base, it makes sense that the pH is above 7. Calculating pH and pOH - High School Chemistry

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Using this value and the pH equation, we can calculate the pH. Now we can find the pOH. The sum of the pH and the pOH is always 14. The pOH of the solution is 7.8. Alternatively, a shortcut can be used to

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estimate the pH. If is in the form , then pH is roughly .

Calculating pH and pOH - High School Chemistry

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